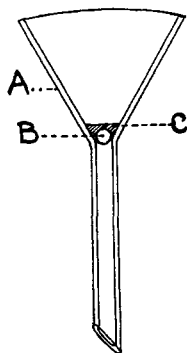


funnel, A, a glass pearl, B, is dropped and over the bead is poured a suspension of finely shredded asbestos, C, to form a layer about $\frac{1}{32}$ of an inch thick. The mat is allowed to drain and suction is applied (see Fig. 1).



A—one inch glass funnel; B—small glass pearl or bead; C—mat of fine grained asbestos. Drawing is $\frac{3}{4}$ actual size.

Fig. 1.

The precipitate and mother liquor are transferred to the micro-filter and the filtrate is removed by gentle suction. One drop of fluid is sufficient to wash the precipitate on the mat. With intermittent suction a precipitate can easily be washed five to ten times with one cc. of solution.

To remove the precipitate the funnel is inverted and a glass rod is inserted into the stem. Precipitates such as calcium, sodium, potassium, phosphorus, etc., that are to be dissolved before determination may be dissolved either on the mat, or in a separate container after removal together with the mat. In the latter case, a second filtration through the same apparatus gives a solution free of asbestos.

The materials for this micro-filter are at hand in every laboratory. They are inexpensive. During a year of use not a single determination has been lost.

CONTRIBUTION FROM THE LABORATORY
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CORRECTION

Through the work of Paul S. Roller¹ it has come to my attention that the signs of two terms in Equation 10 of my paper on titration² are incorrectly recorded. The term $3K_W K_A^{-2}$ in the coefficient of $(H^+)^3$ should be negative, as should also the last term of the equation. I wish also to confirm the result obtained by Roller by an independent method concerning the limiting strength of acid necessary for appearance of an inflection in titration with a strong base, the values of the ionization constants given in my paper for this case being too large, due to a numerical error, by one power of ten.

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¹ Roller, THIS JOURNAL, 50, 1 (1928).

² Eastman, *ibid.*, 47, 332 (1925).